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Chlorometallate and palladium cluster complexes of wide-span diimine and diamine ligands**

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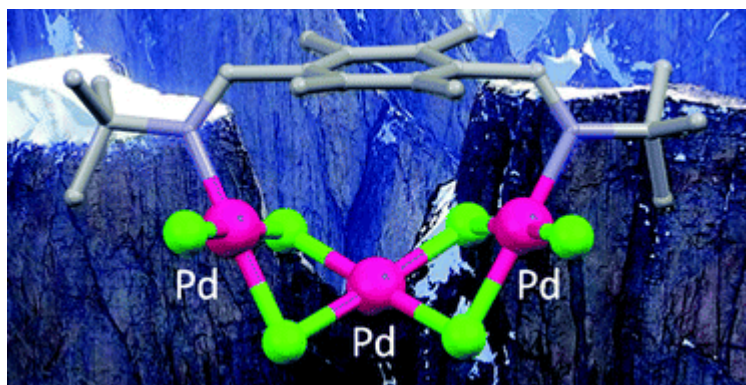
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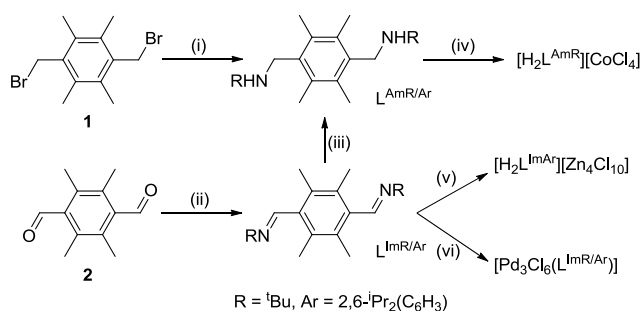
Abstract

Diimine and diamine ligands that are unable to coordinate to a single metal favour the formation of unusual, high-nuclearity Zn chlorometallate and palladium chloride complexes.

Introduction

Ligands that can promote the assembly of multiple metal cations into precisely-structured arrangements are highly desirable as these multimetallic motifs impact on a diverse range of chemical topics such as catalysis,¹ bio-inorganic chemistry,² medicine,³ and single molecule magnetism.⁴ In particular, ligand designs incorporating 1,4-disubstituted arenes have been intensively investigated as platforms that favour bi- and multinuclear complex formation due to the inability of the two donors to chelate to a single metal cation.⁵ For example, imine and amine acyclic ligands,⁶ macrocycles,⁷ and cryptands⁸ derived from 1,4-disubstituted arenes have been developed and show a propensity to form polynuclear complexes, although in simple systems cyclometallation reactions at the central aryl ring are observed, particularly with late transition metals.⁹

We are interested in developing straightforwardly-synthesised and inexpensive ligands for the assembly of multinuclear complexes of metals from across the Periodic Table,¹⁰ in particular to access new chemistry that can occur within well-defined cavities such as found in Pacman complexes,¹¹ and related supramolecular assemblies.¹² We report here on the use of new wide-span diimine and diamine ligands as outer and inner-sphere coordination platforms for the generation of new, high-nuclearity chlorometallate and palladium chloride complexes. The use of durene in the ligand design ensures that undesired cyclometallation reactions are minimised.



Scheme 1. Synthesis of wide-span diimine (L^{Im}) and diamine (L^{Am}) ligands and their chlorometallate and cluster complexes. Reagents and conditions: (i) $\text{H}_2\text{NR}/\text{Ar}$, MeCN; (ii) $\text{H}_2\text{NR}/\text{Ar}$, MeCN; (iii) NaBH_4 , MeOH, $R = {}^t\text{Bu}$; (iv) 2 HCl, Et_2O , CoCl_2 ; (v) 2 HCl, Et_2O , 4 ZnCl_2 ; (vi) 3 $\text{PdCl}_2(\text{MeCN})_2$, CH_2Cl_2 .

The pro-ligands L^{Im} and L^{Am} were synthesised in high yield from the readily available precursors **1** and **2** through straightforward bromide substitution and Schiff-base condensation routes, respectively (Scheme 1). Borohydride reduction of diimine L^{ImR} proved to be the most reproducible route to diamine L^{AmR} . The 2,6-diisopropylaryl-substituted diimine L^{ImAr} and diamine L^{AmAr} compounds were crystallised from saturated solutions of MeCN and their X-ray crystal structures determined (Fig. 1). In the solid state, both L^{ImAr} and L^{AmAr} form extended structures through intermolecular hydrogen-bonding interactions and display no solvent incorporation. In L^{ImAr} (Fig. 1, top), the imine N1-C13 bonds (1.236(2) Å) adopt an anti-conformation (torsion angle = 180.0(2)°) and are offset by 29.8(3)° with respect to the central aryl ring. The sterically-demanding ⁱPr substituents on the terminal aryl groups result in an approximately orthogonal orientation (72.7(2)°) of this aryl ring to the imine N=C group. This orientation results in an intermolecular interaction between the C10 isopropyl methyl hydrogens and the central aryl ring (C10⋯C16 = 3.745(4) Å), so forming a chain motif in the extended structure. Reduction of the N=C double bond in L^{ImAr} to form L^{AmAr} allows free rotation around the N1-C13 single bond (1.489(2) Å) and results in a different structure in the solid state (Fig. 1, bottom) in which the N1-C13 bonds adopt a syn-conformation (torsion angles = 79.9(2)/84.0(2)°). As with L^{ImAr} , the steric demand of the ⁱPr₂(C₆H₃) substituent results in an approximately orthogonal orientation of this group in relation to the amine N1-C13 bond (76.1(2)/68.2(2)°), but in this case the presence of the amino hydrogen results in a hydrogen bond between the two amines (N1⋯N1' = 3.264(2) Å); π-H interactions are also seen in the extended structure.

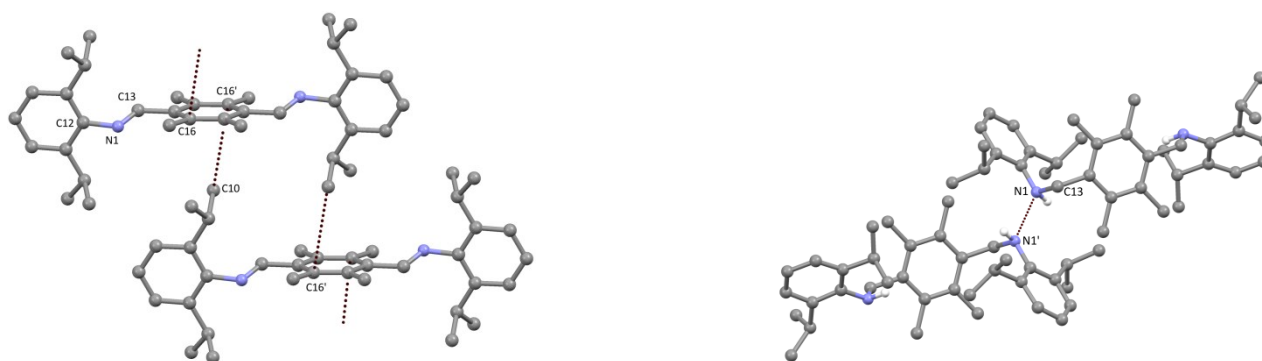


Figure 1. Ball and stick representations of the X-ray crystal structures of diimine L^{ImAr} (left) and diamine L^{AmAr} (right). For clarity, all hydrogen atoms except those involved in hydrogen bonding are omitted.

The coordination chemistry of these ligands was investigated (Scheme 1). Reactions between L^{Im} or L^{Am} and either CoCl_2 or ZnCl_2 in THF showed little change in the resonances for L in the ¹H NMR spectra. However, the use of CDCl_3 as the NMR solvent resulted in the crystallisation of the chlorometallate salts $[\text{H}_2L^{\text{AmR}}][\text{CoCl}_4]$ and $[\text{H}_2L^{\text{ImAr}}][\text{Zn}_4\text{Cl}_{10}]$, presumably as a result of the presence of adventitious HCl in

CDCl₃. Scaled-up reactions were carried out in which the ligands were protonated by HCl (2 eq) in Et₂O prior to the addition of CoCl₂ (1 eq) or ZnCl₂ (4 eq) in Et₂O. This latter procedure resulted in the formation of analytically pure materials. The X-ray crystal structures of both [H₂L^{AmR}][CoCl₄] and [H₂L^{ImAr}][Zn₄Cl₁₀] were determined (Fig. 2) and show that the protonated ligands interact with the chlorometallate salts through a series of hydrogen-bonding interactions; all hydrogens involved in hydrogen-bonding were located in the difference Fourier map and refined with riding thermal parameters and bond distance restraints.

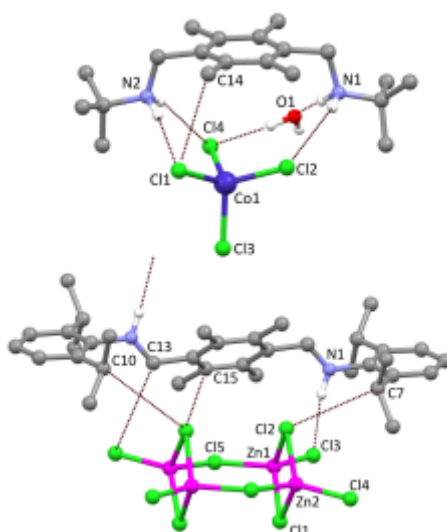


Figure 2. Ball and stick representations of the X-ray crystal structures of [H₂L^{AmR}][CoCl₄](H₂O)(CHCl₃) (top) and [H₂L^{ImAr}][Zn₄Cl₁₀] (bottom). For clarity, CHCl₃ solvent of crystallisation and all hydrogen atoms except those involved in hydrogen bonding are omitted.

In the solid state structure of [H₂L^{AmR}][CoCl₄] (Fig 2., top) the amine appendages adopt a syn-conformation (torsion angle = 6.5(3)^o) that results in a hydrogen-bonded cavity suited for chlorometallate recognition (over chloride), with N-H⋯Cl and C-H⋯Cl hydrogen-bonding interactions ranging in distance from 3.165(3) to 3.357(3) Å. The chelation of the cobaltate anion by the wide-span ligand causes some strain, as evidenced by the 10^o curvature of the central diaminoarene and the inclusion of a molecule of water that interacts both with the protonated amine (N1⋯O1 2.817(5) Å) and the cobaltate (Cl4⋯O1 3.165(3) Å); as with the free ligands above, other intra- and intermolecular close contacts (to CHCl₃ solvent of crystallisation) are present in the extended structure.

In the solid state structure of [H₂L^{ImAr}][Zn₄Cl₁₀] (Fig. 2, bottom), the protonated imine arms adopt an anti-conformation (torsion angle = 176.0(2)^o) and the terminal aryl groups, and to a lesser extent the

central arene, are oriented-orthogonally to the N=C bond (torsion angles = 88.1(3) and 53.4(4)^o, respectively). As with the cobaltate above, a variety of weak interactions exist between the doubly-protonated diimine and the zincate, with N-H \cdots Cl and C-H \cdots Cl hydrogen bonding interactions ranging in distance from 3.152(3) to 3.759(3) Å. However, in this case, the wide-span ligand framework has favoured the isolation of an unusual Zn₄Cl₁₀²⁻ anionic cluster. In this cluster, each zinc cation is tetrahedral with one terminal chloride and three chlorides bridging to adjacent zinc cations. This results in a rectangular arrangement of zinc cations in which two opposing sides of the rectangle comprise either a single bridging chloride or two bridging chlorides, with Zn \cdots Zn distances of 3.8345(5) Å and 3.1405(5) Å and Zn-Cl-Zn angles of 114.25(3) and 85.30(2)^o. Also, the anti-conformation of the imines allows for an intermolecular N1 \cdots Cl3 hydrogen bond (3.152(3) Å) from which a linear chain of alternately sandwiched ligand and metallate clusters grows. While Zn₄Cl₁₀²⁻ anionic clusters have been observed before, to our knowledge they adopt adamantane or open-chain structural motifs and not the new rectangular arrangement of Zn cations seen here.¹³ Furthermore, metal chloride clusters that adopt a similar structural motif are rare, and limited primarily to *o*-phenylene-metallated Pt₄ chlorides,¹⁴ Mo₄Cl₁₂³⁻ clusters,¹⁵ and the chlorobismuthate and antimonates M₄Cl₁₈⁶⁻ (M = Bi, Sb).¹⁶ Significantly, the recognition of high-nuclearity chlorometallates by L^{Im} and L^{Am} may impact on the high-efficiency, hydrometallurgical extraction of base and precious metals.¹⁷

The reaction between PdCl₂(MeCN)₂ and either L^{ImAr} or L^{ImR} in CH₂Cl₂ led to the sole formation of the trinuclear palladium chloride cluster [Pd₃Cl₆(L)]; this formulation was supported by elemental analysis. The X-ray crystal structures of [Pd₃Cl₆(L^{ImAr})] and [Pd₃Cl₆(L^{ImR})] were determined (Fig. 3) and were found to be isostructural; as such, only the former structure will be discussed (Fig.3, top). In a similar manner to [H₂L^{AmR}][CoCl₄], the imine appendages adopt a syn-conformation (torsion angle = 6(2)^o) that allows the ligand to chelate to the single Pd₃Cl₆ cluster through the imine nitrogens N1 and N2. Also, the terminal and central aryl rings are similarly tilted with respect to the N=C bond by 67(4) and 73(4)^o, respectively. Each Pd cation is square planar but the Pd₃Cl₆ cluster is not linear due to the constraints imposed by the chelating diimine ligand. This feature results in a cradle arrangement in which the Pd1-Pd2-Pd3 angle is 127.4(4)^o with Pd(μ-Cl)Pd angles of approximately 90^o.

The solid state structure of [Pd₃Cl₆(L^{ImAr})] is retained in solution. In the ¹H NMR spectrum (Fig. S1, Supplementary information), only one imine resonance is seen at 8.38 ppm, whereas two distinct CH₃ environments are seen at 3.75 and 2.09 ppm for the central arene, which reflects the asymmetry of binding to the Pd₃Cl₆ cluster. Furthermore, free rotation about the N-C bond for the terminal 2,6-diisopropyl aryl substituent is hindered, as reflected by the presence of two isopropyl CH resonances at 3.56 and 3.49 ppm and four resonances for the CH₃ environments. The ¹H NMR spectrum of [Pd₃Cl₆(L^{ImR})] is similar (Fig. S2), indicating that the solid state structure of this complex is also retained in solution, but shows only one resonance for the ^tBu substituent at 1.22 ppm. Furthermore, the ESI mass

spectrum of $[\text{Pd}_3\text{Cl}_6(\text{L}^{\text{ImR}})]$ shows an ion at m/z 682 for the partially-intact cluster $[\text{Pd}_3\text{Cl}_5(\text{L}^{\text{ImR}})]^+$, and indicates structural retention in solution.

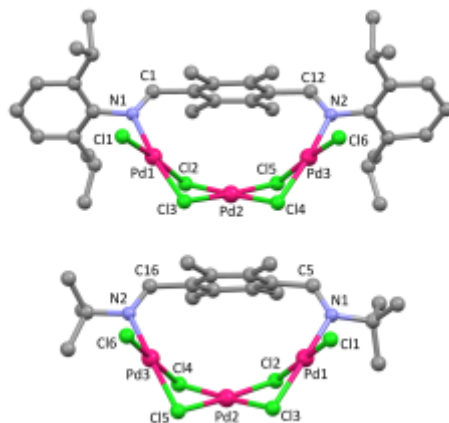


Figure 3. Ball and stick representations of the X-ray crystal structures of $[\text{Pd}_3\text{Cl}_6(\text{L}^{\text{ImAr}})]$ and $[\text{Pd}_3\text{Cl}_6(\text{L}^{\text{ImR}})]$. For clarity, disorder components, solvent of crystallisation, and all hydrogen atoms are omitted.

High nuclearity complexes of Pd chloride are rare and the $[\text{Pd}_3\text{Cl}_6(\text{L})]$ complexes can be viewed as short fragments of polymeric $(\text{PdCl}_2)_n$. While postulated through indirect evidence,¹⁸ only a single example of the crystallographic characterisation of a PdCl_2 trimer end-capped by organic ligands has been described;¹⁹ in this example, the unusual Pd_3Cl_6 motif was stabilised through the use of a large cone-angled, bowl-shaped triarylphosphine. Trinuclear PdCl structures are also known for allyl-complexes, although in these cases one chloride has been substituted by the allyl ligand.²⁰ The related dianion $\text{Pd}_3\text{Cl}_8^{2-}$ has also been characterised structurally as its Bu_4N ammonium salt.²¹ In all of these previous structures, the Pd_3Cl_6 motif is linear, in contrast to the cradle structure seen by us in $[\text{Pd}_3\text{Cl}_6(\text{L})]$. Presumably, this new bent geometry for the Pd_3Cl_6 cluster is necessary to allow a good fit to the wide-span diimine ligands. Furthermore, it is evident that methyl substitution of the central arene ligand has ensured that the formation of cyclometallated products is inhibited.

Notes and references

- [1] M. Delferro and T. J. Marks, *Chem. Rev.*, 2011, **111**, 2450; H. Dau, C. Limberg, T. Reier, M. Risch, S. Roggan and P. Strasser, *ChemCatChem*, 2010, **2**, 724; R. M. Haak, S. J. Wezenberg and A. W. Kleij, *Chem. Commun.*, 2010, **46**, 2713; M. J. Wiester, P. A. Ulmann and C. A. Mirkin, *Angew. Chemie. Int. Ed.*, 2010, **50**, 114.
- [2] J. P. Collman, R. Boulatov, C. J. Sunderland and L. Fu, *Chem. Rev.*, 2004, **104**, 561; R. H. Holm, P. Kennepohl and E. I. Solomon, *Chem. Rev.*, 1996, **96**, 2563.
- [3] Z. Guo and P. J. Sadler, *Angew. Chemie. Int. Ed.*, 1999, **38**, 1512; T. Storr, K. H. Thompson and C. Orvig, *Chem. Soc. Rev.*, 2006, **35**, 534.
- [4] T. Glaser, *Chem. Commun.*, 2011, **47**, 116; G. Aromí and E. K. Brechin, *Struct. Bond.*, 2006, **122**, 1.
- [5] A. L. Gavrilova and B. Bosnich, *Chem. Rev.*, 2004, **104**, 349.
- [6] F. Zhang, M. C. Jennings and R. J. Puddephatt, *Organometallics*, 2004, **23**, 1396.
- [7] S. Brooker and H. L. C. Feltham, *Coord. Chem. Rev.*, 2009, **235**, 1458; P. D. Frischmann and M. J. MacLachlan, *Comments Inorg. Chem.*, 2008, **29**, 26; J. Vela, L. Zhu, C. J. Flaschenriem, W. W. Brennessel, R. J. Lachicotte and P. L. Holland, *Organometallics*, 2007, **26**, 3416; P. Comba, T. W. Hambley, P. Hilfenhaus and D. T. Richens, *J. Chem. Soc., Dalton Trans.*, 1996, 533; P. Comba, A. Fath, T. W. Hambley and D. T. Richens, *Angew. Chemie. Int. Ed.*, 1995, **34**, 1883.
- [8] G. Morgan, V. McKee and J. Nelson, *Inorg. Chem.*, 1994, **33**, 4427; Y. Shu-Yan, L. Qin-Hui, S. Meng-Chang and H. Xiao-Yun, *Inorg. Chem.*, 1994, **33**, 1251.
- [9] Y.-F. Han, H. Li, L.-H. Weng and G.-X. Jin, *Chemical Communications*, 2010, **46**, 3556; B. J. O'Keefe and P. J. Steel, *Organometallics*, 1998, **17**, 3621.
- [10] J. B. Love, *Chem. Commun.*, 2009, 3154
- [11] A. M. J. Devoille, P. Richardson, N. L. Bill, J. L. Sessler and J. B. Love, *Inorg. Chem.*, 2011, **50**, 3116; P. L. Arnold, E. Hollis, F. J. White, N. Magnani, R. Caciuffo and J. B. Love, *Angew. Chem., Int. Ed.*, 2011, **50**, 887; J. W. Leeland, A. M. Z. Slawin and J. B. Love, *Organometallics*, 2010, **29**, 714; E. Askarizadeh, S. B. Yaghoob, D. M. Boghaei, A. M. Z. Slawin and J. B. Love, *Chem. Commun.*, 2010, **46**, 710; P. L. Arnold, A.-F. Pécharman, E. Hollis, A. Yahia, L. Maron, S. Parsons and J. B. Love, *Nat. Chem.*, 2010, **2**, 1056; E. Askarizadeh, A. M. J. Devoille, D. M. Boghaei, A. M. Z. Slawin and J. B. Love, *Inorg. Chem.*, 2009, **48**, 7491; P. L. Arnold, D. Patel, C. Wilson and J. B. Love, *Nature*, 2008, **451**, 315.

- [12] J. W. Leeland, F. J. White and J. B. Love, *J. Am. Chem. Soc.*, 2011, **133**, 7320; J. S. Hart, F. J. White and J. B. Love, *Chem. Commun.*, 2011, **47**, 5711; J. W. Leeland, F. J. White and J. B. Love, *Chem. Commun.*, 2011, **47**, 4132; P. L. Arnold, N. A. Potter, C. D. Carmichael, A. M. Z. Slawin, P. Roussel and J. B. Love, *Chem. Commun.*, 2010, **46**, 1833.
- [13] F. Bottomley, P. D. Boyle, S. Karslioglu and R. C. Thompson, *Organometallics*, 1993, **12**, 4090; F. Bottomley and S. Karslioglu, *J. Chem. Soc., Chem. Commun.*, 1991, 222; J. Estager, P. Nockemann, K. R. Seddon, M. Swadźba-Kwaśny and S. Tyrrell, *Inorg. Chem.*, 2011, **50**, 5258.
- [14] P. M. Cook, L. F. Dahl and D. W. Dickerhoof, *J. Am. Chem. Soc.*, 1972, **94**, 5511.
- [15] B. A. Aufdembrink and R. E. McCarley, *J. Am. Chem. Soc.*, 1986, **108**, 2474.
- [16] N. Leblanc, W. Bi, N. Mercier, P. Auban-Senzier and C. Pasquier, *Inorg. Chem.*, 2010, **49**, 5824; A. Piecha, R. Jakubas, V. Kinzhybalo and T. Lis, *J. Mol. Struct.*, 2008, **887**, 194.
- [17] R. J. Ellis, J. Chartres, K. C. Sole, T. G. Simmance, C. C. Tong, F. J. White, M. Schröder and P. A. Tasker, *Chem. Commun.*, 2009, 583.
- [18] P. M. Maitlis, D. Pollock, M. L. Games and W. J. Pryde, *Can. J. Chem.*, 1965, **43**, 470.
- [19] Y. Ohzu, K. Goto and T. Kawashima, *Angew. Chemie. Int. Ed.*, 2003, **42**, 5714; Y. Ohzu, K. Goto, H. Sato and T. Kawashima, *J. Organomet. Chem.*, 2005, **690**, 4175.
- [20] P. M. Bailey, E. A. Kelley and P. M. Maitlis, *J. Organomet. Chem.*, 1978, **144**, C52; B. T. Donovan, R. P. Hughes, P. P. Spara and A. L. Rheingold, *Organometallics*, 1995, **14**, 489; M. Parra-Hake, M. F. Rettig and R. M. Wing, *Organometallics*, 1983, **2**, 1013.
- [21] S. Schwarz, J. Strähle and U. Weisser, *Z. Anorg. Allg. Chem.*, 2002, **628**, 2495.